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Acids Bases And Salts Chapter

Potassium Chloride (KCl): It is formed after the reaction between potassium hydroxide (a strong base) and hydrochloric acid (a strong acid). (ii) Acidic Salts: Salts which are formed after the reaction between a strong acid and weak base are called Acidic salts. The pH value of acidic salt is lower than 7.

Acids Bases and Salts Class 10 Notes Science Chapter 2

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The reaction between an acid and a base to give a salt and water is known as a neutralisation reaction. In general, a neutralisation reaction can be written as - Base + Acid \rightarrow Salt + Water

2.1.5 Reaction of Metallic Oxides with Acids Activity 2.7 n

Take a small amount of copper oxide in a beaker and add dilute hydrochloric acid slowly while stirring.

CHAPTER2 Acids, Bases and Salts

A salt is a combination of an anion of an acid and a cation of a base. Examples - KCl, NaNO₃, CaSO₄, etc. Salts are usually prepared by the neutralisation reaction of an acid and a base.

Class 10 Chapter 2 Science Notes Acids, Bases, and Salts

Acids - Acids are compounds which give hydrogen ion in water solution. For example, Hydrochloric acid (HCl), Sulphuric acid (H₂SO₄), Nitric acid (HNO₃). Bases - Bases are compounds which give hydroxide ion in water solution. For example, Sodium

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hydroxide (NaOH), Potassium hydroxide (KOH), Calcium hydroxide (Ca (OH) 2).

Chapter Notes: Acids Bases and Salts - Class 10 Science

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Acids, Bases and Salts Notes of CBSE Class 10 Science Chapter with detailed explanation of the chapter 'Acids, bases and salts' along with meanings of difficult words. Given here is the complete explanation of the chapter, along with examples and all the exercises, Question and Answers given at the back of the chapter.

Acids, Bases and Salts Class 10 Science Chapter ...

Chapter - 2, Acids Bases and Salts Notes. • These are the substances which have sour in taste. • They turns blue litmus solution to red. • These substances are bitter in taste. • They turns red litmus solution to blue. • They give OH ions in aqueous

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solution. There are the substance that changes their colour / smell in different types of substances.

Chapter - 2, Acids Bases and Salts Notes

Salts . When acid and base neutralize, salts are formed. Strong acid and strong base combines to form neutral salt. $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$. Eq.1. Formation of Neutral Salt. Strong acid and weak base combine to form acidic salt. For Example, Hydrochloric Acid and ammonium hydroxide combine to form ammonium chloride. Other examples, sodium hydrogen carbonate, sodium hydrogen sulphate etc.

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Answers to know their preparation level.

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Acids, bases and salts are the three important groups of chemical substances that are used by us in different ways. Some of the acids, bases and salts occur in nature and they can be made artificially in factories also. Edible substances and their tastes

Acids, Bases and Salts Class 7 Notes Science Chapter 5 ...

Reaction of Acids and Bases with each other • Acids and Bases react to form salt and water. Acid + Base → Salt + H₂O •

Neutralisation Reaction: Reaction of acid with a base is called as neutralization reaction.

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