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C: When $n = 3$, $l = 0, 1, \text{ or } 2$. The highest value for l in this case is 2, which corresponds to the d subshell. Although the $3d$ block appears to be part of the fourth period, it still has the principal quantum number $n = 3$. In general, the subshells within an energy shell increase in energy as follows: $s < p < d < f$ (although there is no $3f$ subshell).

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