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What is the pH of a .001 M NaOH? pH of a Strong Electrolyte: The pH of an aqueous solution is defined as the negative logarithm of the hydrogen or hydronium ion concentration. The pH scale ranges...

What is the pH of a .001 M NaOH? | Study.com

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Calculate the pH of 0.001m naoh?

Answer. Top Answer. Wiki User.

2010-12-16 20:41:20. 2010-12-16

20:41:20. $-\log(0.001 \text{ NaOH}) = 3$

subtracted from 14. = 11 pH.

Calculate the pH of 0.001m naoh - Answers

What is the ph of 0.001 m naoh ?? -

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19641553

What is the ph of 0.001 m naoh ?? - Brainly.in

The pH of a `0.001 M` aqueous solution of sodium hydroxide will be.

The pH of a `0.001 M` aqueous solution of sodium hydroxide will be

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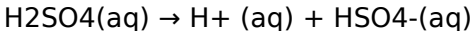
Using a pH indicator strip will tell you that NaOH (sodium hydroxide) is a strong alkaline. This means it has a pH toward the top end of the pH scale, which ranges from 0 to 14. To calculate the exact pH, work out the molarity of the solution, then apply that to the formula for pH.

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How to Calculate the PH of NaOH | Sciencing

I do not agree with the answers that are based on the assumption that H₂SO₄ produces 2 x H⁺ ions per formula unit .

The first dissociation of H₂SO₄ is:



Because this dissociation is total , H⁺] = 0.001 M in this ques...

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What is the pH of 0.001 M H₂SO₄? - Quora

Adding NaOH will increase the pH of water, because NaOH is a base. At 25°C: pH < 7 is an acidic solution pH = 7 is a neutral solution pH > 7 is a basic solution

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How do you calculate the pH of a 0.001 M NaOH solution ...

What is the pH of a .001 M NaOH solution? After adding 21.7 mL of 0.20 M NaOH, the pH is 5.79.

What is the pH of 0.001 N NaOH?

Calculate The PH And POH Of A 0.001 M Solution Of Sodium Hydroxide (NaOH) At

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25 Degree C. Question: Calculate The PH And POH Of A 0.001 M Solution Of Sodium Hydroxide (NaOH) At 25 Degree C. This problem has been solved!

Solved: Calculate The PH And POH Of A 0.001 M Solution Of ...

For 0.001 M NaOH: $[\text{OH}^-] = 0.001 \text{ M}$. \therefore
 $\text{pOH} = -\log [\text{OH}^-] \therefore \text{pOH} = -\log (0.001$

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M) = 3.

what is the ph of 0.001 m NaOH - Brainly.com

i) $+0.001 \text{ M HNO}_3$ $[\text{H}^+] = 0.001 \text{ M} \therefore \text{pH} = -\log(0.001) = 3.0$ ii) 0.001 M NaOH $[\text{OH}^-] = 0.001 \text{ M} \therefore \text{pOH} = -\log(0.001) = 3.0$ and $\text{pH} = 14.00 - \text{pOH} = 14.00 - 3.0 = 11.0$ iii) The solution resulting from

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mixing 400 mL of 0.05 M HCl with 600 mL of 0.05 M NaOH. There are a number of ways to solve this question; here is one way.

From page 35, 38 and 39 of lecture notes Question with answers

$0.001 \text{ M NaOH} \times 1 \text{ mol OH}^- = 1 \text{ mol NaOH}$
 $= 0.001 \text{ M OH}^- \quad 0.001 \text{ M NaOH} \times 1 \text{ m}$

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o l O H – 1 m o l N a O H = 0.001 M O H
– Since $\text{Ba}(\text{OH})_2$ B a (O H) 2 is more basic with a higher hydroxide ion concentration,...

Which solution has the higher pH, a 0.001 M solution of ...

A student is provided with a stock solution of NaOH with 0.01 M

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concentration. A 50.0 mL of stock solution of NaOH is diluted with water to prepare 500.0 mL NaOH solution. a. Calculate the hydronium ion concentration, $[H_3O^+]$ in the final solution. b. Calculate the pH of the diluted NaOH solution. c. Is the final diluted solution acidic, basic or ...

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